



Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CANDIDATE
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CHEMISTRY

0620/43

Paper 4 Theory (Extended)

October/November 2017

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use an HB pencil for any diagrams or graphs.

Do not use staples, paper clips, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES.

Answer **all** questions.

Electronic calculators may be used.

A copy of the Periodic Table is printed on page 12.

You may lose marks if you do not show your working or if you do not use appropriate units.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **12** printed pages.



2

1 Substances can be classified as elements, compounds or mixtures.

State whether each of the following is an element, a compound or a mixture.

(a) brass [1]

(b) gold [1]

(c) butane [1]

(d) air [1]

[Total: 4]

2 (a) (i) Define the term *molecule*.

.....
 [2]

(ii) Define the term *element*.

.....
 [1]

(b) The table shows the composition of four atoms or ions, **A**, **B**, **C** and **D**.

	number of protons	number of neutrons	number of electrons
A	10	10	10
B	10	12	10
C	12	10	10
D	13	14	10

(i) What is the atomic number of **A**?

..... [1]

(ii) What is the nucleon number of **B**?

..... [1]

(iii) Which of **A**, **B**, **C** and **D** are isotopes of each other?

..... [1]

(iv) Which of **A**, **B**, **C** and **D** are atoms?

..... [1]

(v) Which of **A**, **B**, **C** and **D** are positive ions?

..... [1]

(c) Complete the table.

	number of protons	number of electrons
Na		
S ²⁻		
Cl ₂		

[3]

[Total: 11]

(c) (i) Describe the bonding in iron. Include a diagram in your answer.

.....
..... [3]

(ii) Use your diagram in **(c)(i)** to explain why iron is malleable.

.....
..... [2]

(iii) Iron containing a small amount of carbon is known as steel.

Explain why steel is less malleable than iron.

.....
..... [2]

(d) (i) When iron is added to dilute sulfuric acid, an aqueous solution of iron(II) sulfate is formed as one of the products.

Write a chemical equation for the reaction.

..... [1]

(ii) When iron(III) oxide is added to dilute sulfuric acid, an aqueous solution of iron(III) sulfate is formed as one of the products.

Write a chemical equation for the reaction.

..... [3]

6

(e) Aqueous sodium hydroxide, aqueous potassium iodide and aqueous acidified potassium manganate(VII) are added to aqueous solutions of iron(II) sulfate and iron(III) sulfate.

- Iron(II) ions, Fe^{2+} , are reducing agents in aqueous solution.
- Iron(III) ions, Fe^{3+} , are oxidising agents in aqueous solution.

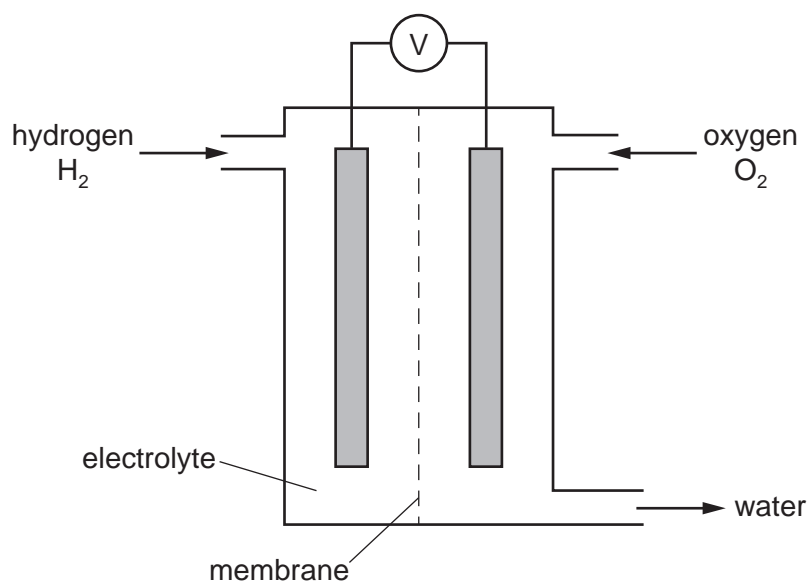
Complete the table.

reagent	observations with aqueous iron(II) sulfate	observations with aqueous iron(III) sulfate
aqueous sodium hydroxide	green precipitate	
aqueous potassium iodide		
aqueous acidified potassium manganate(VII)		no change

[4]

[Total: 22]

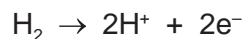
4 Hydrogen and oxygen react together in a hydrogen fuel cell. A hydrogen fuel cell is shown in the diagram.



(a) Name the process by which oxygen is obtained from air.

..... [1]

- (b) (i) In a hydrogen fuel cell, the hydrogen molecules are converted into hydrogen ions, H^+ , according to the ionic half-equation shown.



What type of reaction does this ionic half-equation represent?

..... [1]

- (ii) What **type** of substance reacts by donating hydrogen ions, H^+ ?

..... [1]

- (c) Write a chemical equation for the overall reaction that occurs in a hydrogen fuel cell.

..... [1]

- (d) Hydrogen fuel cells are being developed as alternatives to petrol engines in cars.

- (i) Give **one** advantage of hydrogen fuel cells compared to petrol engines.

..... [1]

- (ii) Give **one** disadvantage of hydrogen fuel cells compared to petrol engines.

..... [1]

- (e) Some fuel cells use ethanol, C_2H_5OH , instead of hydrogen. Carbon dioxide and water are products of the reaction in an ethanol fuel cell.

- (i) Write a chemical equation for the overall reaction occurring in an ethanol fuel cell.

..... [2]

- (ii) State an environmental problem caused by the release of carbon dioxide into the atmosphere.

..... [1]

- (iii) Name the process by which ethanol can be manufactured from a renewable resource.

..... [1]

- (f) Name the process occurring when electrical energy is used to break down an ionic compound.

..... [1]

[Total: 11]

- 5 (a) (i) Name the products formed when sodium nitrate is heated.

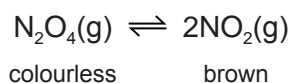
..... [2]

- (ii) When copper(II) nitrate, $\text{Cu}(\text{NO}_3)_2$, undergoes thermal decomposition, three products are formed. One of the products is nitrogen dioxide, NO_2 .

Write a chemical equation for the thermal decomposition of copper(II) nitrate.

..... [2]

- (b) The chemical equation shows the equilibrium between dinitrogen tetroxide (N_2O_4 , a colourless gas) and nitrogen dioxide (NO_2 , a brown gas).



A mixture of dinitrogen tetroxide and nitrogen dioxide is allowed to reach equilibrium in a closed gas syringe.

- (i) In chemistry, what is meant by the term *equilibrium*?

.....

 [2]

- (ii) If the equilibrium mixture is heated at constant pressure, a darker brown colour is seen inside the gas syringe.

What does this information indicate about the decomposition of dinitrogen tetroxide? Explain your answer in terms of the position of the equilibrium.

.....

 [2]

- (iii) Suggest what you would see if the pressure on the equilibrium mixture were increased at constant temperature. Explain your answer in terms of the position of the equilibrium.

.....

 [2]

[Total: 10]

6 (a) Alkanes and alkenes are two homologous series of hydrocarbons.

(i) What is meant by the term *hydrocarbon*?

..... [1]

(ii) What is the general formula of the homologous series of

alkanes,

alkenes?

[2]

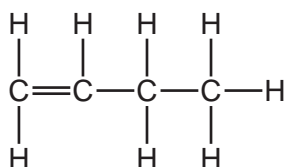
(iii) Other than having a general formula, state **two** characteristics of a homologous series.

1

2

[2]

(iv) The structure of an alkene molecule with the molecular formula C_4H_8 is shown.



Draw the structure of a different alkene molecule with the molecular formula C_4H_8 . Show all of the atoms and all of the bonds.

[1]

(v) What term describes molecules with the same molecular formula but different structural formulae?

..... [1]

- (b)** 25 cm³ of a gaseous hydrocarbon, C_xH_y, were burnt in 150 cm³ of oxygen. This was an excess of oxygen.

After cooling, the volume of the gases remaining was 100 cm³. This consisted of 75 cm³ of carbon dioxide and 25 cm³ of unreacted oxygen. The water that was produced in the reaction was liquid.

All volumes were measured at the same temperature and pressure.

- (i)** What is meant by an excess of oxygen?

..... [1]

- (ii)** What was the volume of oxygen that reacted with the hydrocarbon?

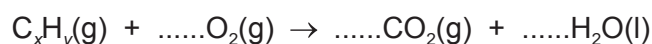
..... cm³ [1]

- (iii)** Complete the table to show the smallest whole number ratio of volumes.

	volume of hydrocarbon reacted	:	volume of oxygen reacted	:	volume of carbon dioxide produced
smallest whole number ratio of volumes		:		:	

[1]

- (iv)** Use your answer to **(b)(iii)** to balance the chemical equation. Deduce the formula of the hydrocarbon.



formula of the hydrocarbon = [2]

[Total: 12]

- 7 (a) Carbon and silicon are elements in Group IV of the Periodic Table.
Carbon dioxide from the air moves into green plants and is converted into carbohydrates.
- (i) Name the process by which carbon dioxide molecules move through the air into green plants.
..... [1]
- (ii) Explain why silicon(IV) oxide **cannot** move through the air in the same way that carbon dioxide can.
..... [1]
- (iii) Name the process by which carbon dioxide is converted into glucose, $C_6H_{12}O_6$, in green plants. Give **two** conditions required for this process to occur. Write a chemical equation for the reaction which occurs.
- name of process
- condition 1
- condition 2
- chemical equation [5]

(b) Starch is a natural polymer made from glucose.

- (i) What type of polymerisation occurs when glucose is converted into starch?
..... [1]
- (ii) What type of reaction occurs when starch is converted into glucose?
..... [1]
- (iii) Starch can be represented as shown.



Complete the diagram below to represent the structure of the glucose monomer.



[1]

[Total: 10]

The Periodic Table of Elements

		Group							
I	II	III	IV	V	VI	VII	VIII		
1	2	3	4	5	6	7	8	9	10
H hydrogen 1	He helium 4	B boron 11	C carbon 12	N nitrogen 14	O oxygen 16	F fluorine 19	Ne neon 20		
Key									
atomic number									
atomic symbol									
name									
relative atomic mass									
3	4	5	6	7	8	9	10	11	12
Li lithium 7	Be beryllium 9	B boron 11	C carbon 12	N nitrogen 14	O oxygen 16	F fluorine 19	Ne neon 20		
11	12	13	14	15	16	17	18		
Na sodium 23	Mg magnesium 24	Al aluminium 27	Si silicon 28	P phosphorus 31	S sulfur 32	Cl chlorine 35.5	Ar argon 40		
19	20	21	22	23	24	25	26	27	28
K potassium 39	Ca calcium 40	Sc scandium 45	Ti titanium 48	V vanadium 51	Cr chromium 52	Mn manganese 55	Fe iron 56	Co cobalt 59	Ni nickel 59
37	38	39	40	41	42	43	44	45	46
Rb rubidium 85	Sr strontium 88	Y yttrium 89	Zr zirconium 91	Nb niobium 93	Mo molybdenum 96	Tc technetium —	Ru ruthenium 101	Rh rhodium 103	Pd palladium 106
55	56	57–71	72	73	74	75	76	77	78
Cs caesium 133	Ba barium 137	lanthanoids	Hf hafnium 178	Ta tantalum 181	W tungsten 184	Re rhenium 186	Os osmium 190	Ir iridium 192	Pt platinum 195
87	88	89–103	104	105	106	107	108	109	110
Fr francium —	Ra radium —	actinoids	Rf rutherfordium —	Db dubnium —	Sg seaborgium —	Bh bohrium —	Hs hassium —	Mt meitnerium —	Ds darmstadtium —
81	82	83	84	85	86	87	88	89	90
Tl thallium 204	Pb lead 207	Bi bismuth 209	Po polonium —	At astatine —	Rn radon —	Ac actinium —	Th thorium 232	Pa protactinium 231	U uranium 238
91	92	93	94	95	96	97	98	99	100
Pr praseodymium 141	Nd neodymium 144	Pm promethium —	Sm samarium 150	Eu europium 152	Gd gadolinium 157	Tb terbium 159	Dy dysprosium 163	Ho holmium 165	Er erbium 167
57	58	59	60	61	62	63	64	65	66
La lanthanum 139	Ce cerium 140	Pr praseodymium 141	Nd neodymium 144	Pm promethium —	Sm samarium 150	Eu europium 152	Gd gadolinium 157	Tb terbium 159	Dy dysprosium 163
89	90	91	92	93	94	95	96	97	98
Ac actinium —	Th thorium 232	Pa protactinium 231	U uranium 238	Np neptunium —	Pu plutonium —	Am americium —	Cm curium —	Bk berkelium —	Cf californium —
67	68	69	70	71	72	73	74	75	76
Ho holmium 165	Er erbium 167	Tm thulium 169	Yb ytterbium 173	Lu lutetium 175	La lanthanum 139	Ce cerium 140	Pr praseodymium 141	Nd neodymium 144	Pm promethium —
101	102	103	104	105	106	107	108	109	110
Md mendelevium —	No nobelium —	Lr lawrencium —	Ac actinium —	Th thorium 232	Pa protactinium 231	U uranium 238	Np neptunium —	Pu plutonium —	Am americium —
111	112	113	114	115	116	117	118	119	120
Rg roentgenium —	Cn copernicium —	Nh nihonium —	Fl flerovium —	Mc moscovium —	Lv livermorium —	Ts tennessine —	Og oganeson —	Uu ununium —	Uub ununbium —

lanthanoids

actinoids

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).